



Marine carbonyl sulfide (OCS) and carbon disulfide (CS₂): a compilation of measurements in seawater and the marine boundary layer

5 Sinikka T. Lennartz^{1,*}, Christa A. Marandino¹, Marc von Hobe², Meinrat O. Andreae^{3,4}, Kazushi Aranami⁵, Elliot Atlas⁶, Max Berkelhammer⁷, Heinz Bingemer⁸, Dennis Booge¹, Gregory Cutter⁹, Pau Cortes¹⁰, Stefanie Kremser¹¹, Cliff Law¹², Andrew Marriner¹², Rafel Simó¹⁰, Birgit Quack¹, Günther Uher¹³, Huixiang Xie¹⁴, Xiaobin Xu¹⁵

10 ¹GEOMAR Helmholtz-Centre for Ocean Research Kiel, Marine Biogeochemistry, Düsternbrooker Weg 20, 24105 Kiel,
* now at: University of Oldenburg, Institute for Chemistry and Biology of the Marine Environment, Marine Geochemistry,
Carl-von-Ossietzky-Straße 9-11, Oldenburg

²Forschungszentrum Jülich GmbH, IEK-7, Jülich, Germany

³Max Planck Institute for Chemistry, P.O. Box 3060, D-55020 Mainz, Germany

⁴Scripps Institution of Oceanography, University of California San Diego, La Jolla, CA 92093-0230, USA

15 ⁵Hokkaido University, Japan

⁶University of Miami, Atmospheric Sciences, Miami, FL, USA

⁷University of Illinois at Chicago, Dept. of Earth and Environmental Sciences, Chicago, IL, USA

⁸Goethe-University Frankfurt am Main, Institut für Atmosphäre und Umwelt, Frankfurt, Germany

⁹Old Dominion University, dept. Ocean, Earth & Atmospheric Sciences, Norfolk, VA, USA

20 ¹⁰Institut de Ciències del Mar, Departament de Biologia Marina i Oceanografia, Barcelona, Spain

¹¹Bodeker Scientific, Alexandra, New Zealand

¹²National Institute of Water and Atmospheric Research (NIWA), Wellington, New Zealand

¹³School of Natural and Environmental Science, Newcastle University, Newcastle upon Tyne, U. K.

¹⁴Institut des sciences de la mer de Rimouski, Université du Québec à Rimouski, Rimouski, Canada

25 ¹⁵Key Laboratory for Atmospheric Chemistry of China Meteorology Administration, Chinese Academy of Meteorological Sciences, Zhongguancun Nandajie 46, Beijing 100081, China

Correspondence to: Sinikka T. Lennartz (sinikka.lennartz@uni-oldenburg.de)

30 **Abstract.** Carbonyl sulfide (OCS) and carbon disulfide (CS₂) are volatile sulfur gases that are naturally formed in seawater and exchanged with the atmosphere. OCS is the most abundant sulfur gas in the atmosphere, and CS₂ is its most important precursor. They have gained interest due to their direct (OCS) or indirect (CS₂ via oxidation to OCS) contribution to the stratospheric sulfate aerosol layer. Furthermore, OCS serves as a proxy to constrain terrestrial CO₂ uptake by vegetation. Oceanic emissions of both gases contribute a major part to their atmospheric concentration. Here we present a database of
35 previously published and unpublished, mainly ship-borne measurements in seawater and the marine boundary layer for both gases, available at <https://doi.pangaea.de/10.1594/PANGAEA.905430> (Lennartz et al., 2019). The database contains original measurements as well as data digitalized from figures in publications from 42 measurement campaigns, i.e. cruises



or time series stations, ranging from 1982 to 2019. OCS data cover all ocean basins except for the Arctic Ocean, as well as all months of the year, while the CS₂ dataset shows large gaps in spatial and temporal coverage. Concentrations are consistent across different sampling and analysis techniques for OCS. The database is intended to support the identification of global spatial and temporal patterns and to facilitate the evaluation of model simulations.

5 1 Introduction

Carbonyl sulfide (OCS) is the most abundant sulfur gas in the atmosphere with a tropospheric mixing ratio around 500 ppt (Kremser et al., 2016). Carbon disulfide (CS₂) is a short-lived sulfur gas, which is oxidized within hours to days. Because OCS is a major product of this oxidation with a yield of 82 % (i.e. 82 molecules of OCS produced from 100 CS₂ molecules), CS₂ oxidation is a major source of OCS in the atmosphere.

10 The atmospheric mixing ratio of OCS has been stable for the past four decades and the budget is thus considered balanced (Montzka et al., 2007; Kremser et al., 2016). Due to its long tropospheric lifetime of 2-7 years, OCS is entrained into the stratosphere. In volcanically quiescent periods, OCS (and indirectly CS₂) is thought to be a major contributor to stratospheric sulfate aerosols that influence the radiative budget of the Earth (Crutzen, 1976; Brühl et al., 2012). In addition, OCS can be used as a proxy to quantify the CO₂ uptake of plants (gross primary production), which is a major source of uncertainty in
15 climate modelling (Whelan et al., 2018). Both scientific interests benefit from a well constrained atmospheric budget. OCS and CS₂ are produced naturally in the ocean, and their oceanic emissions contribute substantially to their atmospheric concentrations (Chin and Davis, 1993; Watts, 2000; Kremser et al., 2015).

Oceanic source estimates of OCS and its precursor CS₂ still contain large uncertainties (Kremser et al., 2016; Whelan et al., 2018). Current efforts to model surface concentrations of OCS in seawater diverge in their results (Launois et al., 2015;
20 Lennartz et al., 2017). Most measurements on oceanic OCS and CS₂ were performed in the 1980s and 1990s, and data are often not available or stored inaccessibly for use, hampering model evaluation or analysis of global spatial and temporal concentration patterns. Therefore, a combined database for marine measurements of OCS and CS₂ has been given high priority in a recent review on using OCS as a tracer for gross primary production (Whelan et al. 2018). Here, we aim to provide such a comprehensive database by compiling previously reported as well as unpublished data, from corresponding
25 authors of the original studies or via digitalization from pdf documents.

Both OCS and CS₂ show a pronounced variability in seawater which implies a need for highly resolved observations. Therefore, we pay special attention to the temporal resolution of measurements in the database. The temporal variability is a direct result of the marine cycling of both gases. Photochemical reactions involving chromophoric dissolved organic matter (CDOM) lead to the formation of OCS, as does a light-independent production pathway (Ferek and Andreae, 1984; Weiss et al., 1995a; Von Hobe et al., 2001). OCS is efficiently hydrolyzed in seawater. The temperature dependence of the hydrolysis
30 reaction leads to high degradation rates in warm waters (Elliott et al., 1989; Radford-Knoery and Cutter, 1994; Kamysny et al., 2003). The efficient photochemical production, as well as the fast degradation in warm waters result in strong diurnal and



seasonal cycles of OCS in the surface ocean (Kettle et al., 2001; Ulshöfer et al., 1995). CS₂ is photochemically produced in seawater as well, but diurnal cycles are not as pronounced due to lower efficiency of the sink processes. Concentrations of CS₂ and OCS in seawater differ strongly depending on the time of day and season measured. To facilitate interpretation of concentration measurements on larger scales in relation to the processes described above, ancillary data coinciding with trace gas measurements are also reported if available, such as meteorological or physical seawater properties. The database is described with respect to number of data, range and patterns of concentrations, analytical methods, temporal and spatial coverage and sampling frequency for each dataset.

2 Methods

2.1 Data collection

Data were obtained either from authors of previous studies directly or digitalized with a web based digitalization tool from pdf documents. Web of Science was searched for the key words ‘carbonyl sulfide’ (both sulfide and sulphide), ‘carbon disulfide’ (both sulfide and sulphide) in connection with ‘ocean’ or ‘seawater’. When data could not be obtained directly from authors, relevant figures were identified and digitalized with the WebPlotDigitizer Automeris (URL: <https://apps.automeris.io/wpd/>, accessed January 2019). When digitalizing the data from documents, concentration data were rounded to the integer to account for uncertainty in the digitalization method introduced, e.g. by misalignment of the axes in case of old, scanned pdf documents. We include here only shipborne measurements or observations from stations with a marine signal (i.e. research platform in the North Sea ID15, Amsterdam Island ID6, Bermuda ID39). For atmospheric OCS data from aircraft campaigns or continental time series stations, i.e. HIAPER-Pole-to-Pole-Observations (HIPPO, Montzka (2013)), Atmospheric Tomography Mission (ATom, Wofsy et al. (2018)) as well as the NOAA time series stations from the Earth System Research Laboratory – Global Monitoring Division (NOAA-ESRL, Montzka (2004), Montzka et al. (2007)) we refer to the respective repositories accessible online (HIPPO: https://www.eol.ucar.edu/field_projects/hippo; ATom: <https://espo.nasa.gov/atom/content/ATom>, NOAA-ESRL: <https://www.esrl.noaa.gov/gmd/>).

Concentration data were converted to the unit picomole OCS/CS₂ per liter, accounting for molar masses of sulfur (32.1 g), OCS (60 g) and CS₂ (76.1 g). Data were collected together with the following metadata (if reported in the original publication or otherwise available):

- Latitude of measurement
- Longitude of measurement
- Date, including year, month, day, hour, minute
- Name of the cruise and/or ship
- Contributor
- Main reference for data
- Method description
- Main reference for method



- Sample depth
- Any ancillary data (meteorological, physical, biological data)
- Flag describing the sampling resolution (see Tab. 1).

It should to be noted that several commonly used materials such as any rubber parts may lead to contaminations when measuring OCS and CS₂. A non-exhaustive list of problematic materials is available here <http://www.cosanova.org/materials-to-avoid.html> (accessed February, 2019). We paid attention to the method description of each dataset, and data were only included when blank measurements are reported or the description of the material was provided (e.g. Teflon used). A filling value of -999 was introduced for concentrations below the respective detection limit of each individual dataset. The database can be accessed at the data repository PANGAEA (<https://doi.pangaea.de/10.1594/PANGAEA.905430>, (Lennartz et al., 2019)).

2.2 Trace gas analysis

2.2.1 Carbonyl sulfide in seawater

Carbonyl sulfide (OCS) concentrations in seawater were commonly measured with a method to separate gaseous OCS from the seawater, connected to a detection system. Two main principles were applied to separate OCS from seawater: 1) purging the water sample with an OCS-free gas to transfer the total dissolved OCS into the gas phase, or 2) using an equilibrator, where a gas phase is brought into equilibrium with the seawater sample. The OCS concentration in water is then calculated using Henry's law and the temperature during the equilibration process. Sampling using method 1 is usually performed discretely, and has sometimes been replaced by method 2 with automated (semi-)continuous sampling with a sampling resolution of <15 minutes since 2015 (Ulshöfer et al., 1995; Von Hobe et al., 2008; Lennartz et al., 2017). OCS detection in discrete samples used gas chromatography (GC). Most GCs were then coupled to a flame photometric detector (GC-FPD), or, less frequently, to an electron capture detector (GC-ECD). Commonly, samples were cryogenically pre-concentrated (i.e. with liquid N₂) prior to injection into the GC. A new technique using off-axis integrated cavity output spectroscopy (OA-ICOS) has only recently been developed to continuously measure dissolved OCS in seawater with the use of an equilibrator (Lennartz et al., 2017).

For the majority of the samples in the database, the precision was reported to be better than 10%, and the limit of detection is around 2 pmol L⁻¹ (see Tab. 2 for details on individual datasets). The instability of OCS in water makes the comparison with liquid standards difficult, which is why most of the studies used permeation tubes to calibrate their instruments. Unfortunately, no inter-calibration between cruises is reported (see section 3.1 for a discussion of quality control of the data).



2.2.2 Carbonyl sulfide in the marine boundary layer

Quantifying the OCS concentration in the sampled gas is performed in a similar way, with the same analytical systems, as described in section 2.2.1 for dissolved concentration measurements. The database consists mainly of shipborne measurements, but includes measurements from two land-based stations with strong marine influence. These two datasets
5 are (1) ID6 from Amsterdam Island in the Southern Ocean, and (2) ID39 from Tudor Hill, Bermuda.

The majority of studies used a GC-FPD system; GC-ECD and OA-ICOS were less frequently used. Detection limits and precision was comparable or identical to seawater measurements described in the previous section. Details on each individual method are listed in Tab. 2. Quantification was achieved with standards produced from permeation tubes and from gas cylinders from various manufacturers (see Tab. 2 for specifications of individual studies). No inter-calibration for the
10 complete database is available.

2.2.3 Carbon disulfide in seawater

Sampling of carbon disulfide (CS₂) was performed discretely from both continuously pumped water and from containers such as Niskin bottles. Concentrations of CS₂ were measured with a sampling frequency of up to 15 minutes. Most
15 frequently, a GC-MS system was used; GC-FPD was less common. Prior to the injection in the GC, samples were either purged with CS₂-free gas, and a cooled trap was used for preconcentration (purge+trap system), or the gas and liquid phase were brought to equilibrium with an equilibrator. Detection limits ranged down to 1 pmol L⁻¹, and the precision was around 3-5% (see Tab. 2 for specification of the individual datasets). Standard measurements include permeation tubes or liquid standards prepared in ethylene glycol, but no inter-calibration has been reported.

20

2.2.4 Carbon disulfide in the marine boundary layer

Samples were commonly taken discretely from the vessel's deck, directly into air canisters, or sampled directly from air drawn with tubing into the laboratory. Altitudes where samples were taken ranged from 10-25m. The detection of CS₂ in the gas phase was similar to the analytical methods described in the section 2.2.3, without the step of purging the gas out of the
25 water. GC-MS or GC-FPD were used for detection. As described above, detection limits range down to 1 ppt, and the precision is ~3-5% (see Tab. 2 for specification of the individual datasets). Standards were either permeation tubes or gas cylinders, detailed in Tab. 2.

3 Description of dataset

All data sets included here provide some means of quality control, and calibration procedures including primary gravimetric
30 standards from permeation tubes or by certified gas standards are described in the original publications. However, the



database compiled here includes measurements made by different laboratories and, thus, different measurement systems. One limitation of the database is the missing intercalibration across these different measurement systems. Since many of these systems were built and deployed in the 1990s and do no longer exist, such an intercalibration is not possible anymore. We strongly recommend undertaking efforts for intercalibration across laboratories for future oceanic measurements of OCS and CS₂. Since no practical quality control is possible, we assess the quality of the database by its internal consistency.

3.1 Carbonyl sulfide in seawater

Measurements of OCS in seawater were collected from 32 cruises, resulting in 7536 individual measurements (Fig. 1, Tab. 2 and 3). OCS concentrations were measured in the picomolar range in the surface and subsurface ocean, with a mean concentration of 32.3 pmol L⁻¹ (n=7536, Fig. 2a), ranging from below the detection limit to 1466 pmol L⁻¹ in Rhode Island river estuary.

The majority of measurements were made in the Atlantic Ocean, least from the Indian Ocean. No measurements are available for the Arctic Sea. The sampling was heavily biased towards surface ocean measurements shallower than 10m depths, and only few measurements (<3%) were obtained from concentration profiles in the water column (144 measurements). The available profiles range down to a water depth of 2000m (Tab. 3). Reporting the sampling depth is critical for photochemically produced substances such as OCS, as the penetration depths of UV light and hence the photoproduction varies spatially and temporally. Samples that were obtained from depths shallower than 10m are referred to as 'surface samples', most of them were obtained at a depth of 3-5m. Maintaining a continuous water supply despite water level changes by waves on a moving vessel is a challenge and currently hinders continuous sampling at shallower depths. Profile measurements from the North Atlantic indicated that differences in concentration within the surface of the mixed layer <10m are in a range of about 5 pmol L⁻¹ (Cutter et al., 2004), but might potentially become larger with surface stratification and high irradiation (Fischer et al., 2018).

OCS measurements were reported in 12 minutely to monthly resolution (Fig. 3e). Hence, the majority of the database has the required temporal resolution to cover the full range of the diurnal variation, i.e. a measurement interval of 4h or less is needed to minimize averaging errors due to interpolation.

The global variability of the available measurements shows lowest concentrations in tropical and subtropical waters, especially compared to concentrations at higher latitudes of the Southern hemisphere (Fig. 3a). The pattern of highest OCS concentrations in coastal and shelf regions has been reported for individual datasets (Cutter and Radford-Knoery, 1993), but is also recognizable in this global database. Especially the data from cruise ID10 (Cutter and Radford-Knoery, 1993), which covers estuaries and shelves, is 10-1000 fold higher than in oligotrophic warm waters (Fig. 4a). This pattern is also evident in elevated concentration in the North Sea (Uher and Andreae, 1997), the Mediterranean Sea (Ulshöfer et al., 1996) and the coastal waters of Amsterdam Island in the Indian Ocean (Mihalopoulos et al., 1992). Concentration profiles in the water column show a typical photochemical behaviour and decrease with depth, although subsurface peaks occur occasionally (Von Hobe et al., 1999; Cutter et al., 2004; Lennartz et al., 2017). Despite the still limited size of the database, it already



covers a large part of the global variability, as it includes measurements from a variety of different biogeochemical regimes, i.e. from oligotrophic waters (Cutter et al., 2004; Lennartz et al., 2017; Von Hobe et al., 2001) to higher trophic stages in shelf (Uher and Andreae, 1997), estuary (Cutter and Radford-Knoery, 1993) and upwelling regions (Ferek and Andreae, 1983; Mihalopoulos et al., 1992; Von Hobe et al., 1999; Lennartz et al., 2017).

5 The annual pattern illustrated in Fig. 3b is different for the Southern and Northern hemisphere. Lowest median concentrations in the Southern hemisphere are present during austral winter months, and increase up to five times during austral summer. In the Northern hemisphere, lowest concentrations are present during late boreal summer (A,S,O) and late boreal winter (J,F,M). The range of observed concentrations is similar for both hemispheres. Compared to the spatial pattern in Fig. 3a, the temporal variability in the database is larger than the spatial variability.

10 Figure 4a illustrates the concentration range of each dataset for OCS in seawater. The internal consistency of the database is supported by: (1) the variation of concentration in this database is consistent with the current process understanding, and thus reflects actual variability. The majority of the global measurements (60%) fall into a very narrow range of 8.7-43.0 pmol L⁻¹, and outliers of this 20-80 percentile range are explicable by location or time of measurement. For example, OCS concentrations during cruise ID2, ID10, ID13 and ID19 were much higher than observed by other cruises (Fig. 4), which can

15 be explained by the location of Chesapeake Bay (ID2, ID13), the Petaquamscutt estuary (ID 10) and North Sea (ID19) in shelf areas or close to estuaries, where high CDOM abundance enhances photochemical production and increases concentration. An example for a particularly low concentration is cruise ID18, which took place during winter. The authors refer the low concentration as due to low photoproduction at that time (Ulshöfer et al., 1995); (2) Measurements obtained from cruises that cover a similar temporal and spatial area yield comparable results, such as cruises ID3 and ID4 (Pacific

20 Ocean), cruises ID27 and ID28 (Atlantic transects) or cruises ID20, ID26, and ID32 from the North Atlantic (Fig. 4); (3) Reported OCS concentrations in seawater and the marine boundary layer are consistent across different laboratories and methods. The introduction of new methods, e.g. OA-ICOS (cruiseID 36 and 39) yields results that are comparable to previous measurements using GC-FPD. To facilitate comparison of individual datasets, they are grouped according to the analysis system used in Tab. 2 (capital letters in last table column).

25 3.2 Carbonyl sulfide in the marine boundary layer

The dataset of OCS in the marine boundary layer includes 14291 measurements from 30 cruises (Fig. 3f, Tab. 3). The average mixing ratio in the marine boundary layer is 548.9 (209-1112) ppt. All major ocean basins were covered, except for the Arctic Ocean. The North Atlantic Ocean including the North Sea was sampled most frequently.

30 Sampling of OCS in the marine boundary layer is done either discretely by pumping air in canisters, or continuously by pumping air directly into the detection system (see Tab. 2 for details). Marine boundary layer air was often sampled from the ship's uppermost deck, and reported sampling heights ranged from 10-35m (Tab. 3). Given the relatively stable atmospheric mixing ratios (compared to the strong diel variations of dissolved OCS), a strong gradient in mixing ratios towards the sea surface is not expected. Hence, the database is suited to calculate the concentration gradient across the air-



sea boundary, making it valuable for calculating oceanic emissions. The sampling frequency in individual datasets ranged from intervals shorter than hourly to monthly time series (Fig. 3f). Given the weak diurnal variability compared to the seasonal variability, the reported resolution in all of the individual datasets is sufficient for large scale comparisons.

The global variability of boundary layer OCS mixing ratio is less pronounced and does not show the same spatial pattern as that of dissolved OCS in the surface ocean (Fig. 2b). Ranges of mixing ratios are similar across all latitude bins, with minor variations (Fig. 3b). In several individual datasets, e.g. in Pacific (Weiss et al., 1995b) and Atlantic transects (Xu et al., 2001), mixing ratios in the tropics increase compared to higher latitudes during the respective cruise (Fig. 2b). Highest atmospheric mixing ratios are reported from around Europe (including the Mediterranean) as well as off the Falkland Islands (Fig. 2b). The complete database includes measurements from January to December in the Northern hemisphere, but data for September are missing in the Southern hemisphere. The seasonal variability in atmospheric mixing ratio was less pronounced compared to the variability in seawater OCS concentration, with monthly medians ranging from 439-647 ppt in the Northern and 467-523 ppt in the Southern hemisphere. No clear seasonal pattern was observed in either of the hemispheric datasets. The lack of such a pattern in atmospheric concentrations might result from the limited size of the dataset and the spatial heterogeneity of the sampling locations (i.e. influence of local vegetation sinks or anthropogenic sources of the air mass history).

The internal consistency of the database is of similar quality as described for OCS in seawater. 60% of the data (i.e. between 20 and 80 percentile) fall in a narrow range of 477-621 ppt (Fig. 4b). Some features are present across different datasets and, hence, support the internal consistency of the dataset: For example, locally elevated mixing ratios in tropical latitudes are present in single datasets, and also globally (Fig. 2b and 4b). Elevated atmospheric mixing ratios were reported by several studies providing measurement from around Europe (Fig. 2b and 4b).

3.3 Carbon disulfide in seawater

Measurements of dissolved CS₂ in seawater are reported from 11 cruises (Fig. 1, 1813 measurements), with an average of 15.7 (1.1-376) pmol L⁻¹. Most of the measurements were performed in the Atlantic Ocean, comprising three Atlantic meridional transects (Fig. 2c). No measurements are available from the Arctic and Antarctic waters and the Indian Ocean. The latitudinal variation of CS₂ in seawater was small (Fig. 5a, on average <5 pmol L⁻¹), although individual studies report a general covariation of concentrations and water temperature (Xie and Moore, 1999; Lennartz et al., 2017). Apart from an Atlantic transect with exceptionally high concentrations (Lennartz et al., 2017), concentrations tend to increase towards coastal and upwelling regions (Fig. 2c), but this increase was less pronounced compared to the spatial variability of OCS (Fig. 2a). Temporal variability of CS₂ water concentrations was larger compared to the spatial variability (Fig. 5a and 5c), but did not show a clear seasonal or spatial pattern. Concentrations were comparable in the Northern and Southern hemisphere. Diurnal variability of surface concentrations was present on some, but not all days within individual datasets, likely representing the varying efficiency of the local sink process in the mixed layer. The occurrence of diel cycles calls for a similarly high sampling frequency as suggested for OCS (i.e. more frequently than 4 hourly). However, most of the dataset



comprises a sampling frequency of daily to monthly, and the sampling is biased towards daytime (Fig. 5e). Hence, averaged concentrations might slightly overestimate diel averages.

The database presented here indicates the common range of seawater concentrations and covers several biogeochemical regimes. However, limitations remain, viz.: 1) a general sparsity of measurements, 2) data gaps especially in high latitudes and 3) insufficient sampling frequency to cover full diel variability in many individual data sets.

The limited size of the database for CS₂ in seawater hampers internal data comparison. The majority of the data (between 20 and 80 percentile) falls in the range of 6.1-15.6 pmol L⁻¹. Individual datasets from the Southern Ocean (cruiseID 7, not georeferenced) and from an Atlantic transect (cruiseID 38) show mean concentrations that are considerably higher than observed on other cruises (Fig. 6a). Since data from dataset ID7 represent the only available measurements for this location, based on this database, we cannot determine whether this is an artefact or not. However, we see a similar trend in the OCS data observed by dataset ID7. Also, the low atmospheric mixing ratio measured during this specific cruise speaks against a contamination problem. For the Atlantic transect it is evident that the average concentration is higher compared to the other three Atlantic cruises with a similar cruise track, i.e. dataset ID28, ID29 and ID33. However, the minimum measured concentration in this specific dataset is 7 pmol L⁻¹, which makes a strong contamination unlikely. The atmospheric mixing ratios during cruise ID38 are also lower than those during the other two Atlantic transects, which negates a strong contamination problem. Furthermore, the covariance with temperature is evident in this and in other datasets (Xie and Moore, 1999; Lennartz et al., 2017). CS₂ concentration in dataset ID38 is reported twice daily, one time at 8-10 a.m. and one time at 15-18 a.m. local time. Potentially, the average is misleading in this respect, because it masks potential diel cycles. Daily maxima of cruise ID38 agree with daily maxima of some parts of the other Atlantic transects (ID28,29,33), but not on the majority of days. The minimum values over large parts of the cruise ID38 were higher than those in the cruises ID28, 29 and 33. Potentially, the minima might have been missed by the coarse sampling.

3.2 Carbon disulfide in the marine boundary layer

CS₂ measurements in the marine boundary layer are only available for the Atlantic Ocean from six cruises, i.e. 1036 individual measurements. Atmospheric mixing ratios were on average 42.2 ppt (2.5 to 275.7 ppt, Fig. 2d).

Since this dataset is only comprised of four individual cruises, any perceived pattern in global variation should be taken with caution, as it might rather reflect natural variability or differences between individual laboratories. Spatiotemporal variability will become clearer once more data are available. Almost no preference is given to new measurement locations or times, as any new dataset will help to further constrain spatial and temporal variability of CS₂ concentration. Latitudinal median mixing ratios varied between studies by a factor of 2, but due to the limited size of the dataset, it is currently unclear if this variation is meaningful. However, a CS₂ mixing ratio of 42.2 ppt in the marine boundary layer may become relevant for calculation of oceanic emissions. Commonly, marine boundary layer concentrations of CS₂ are assumed to be 0 due to its short lifetime, which will lead to an overestimation of emissions in case the true mixing ratio is higher, as our database



indicates (at a temperature of 20°C, a salinity of 35 psu, a wind speed of 7 m s⁻¹ and a CS₂ water concentration of 16 pmol L⁻¹, the difference between 0 and 42 ppt CS₂ in air leads to an overestimation of 21%).

Due to the sparsity of data and the expected strong variability resulting from the short atmospheric lifetime, we will not use this limited dataset here for assessing the internal consistency across locations. The variation between the two Atlantic
5 transect datasets ID28 and ID29 with a strong overlap in measured mixing ratios seems reasonable (Fig. 6b), but more data is needed to establish a comparison on larger scale.

4 Recommendations for oceanic OCS and CS₂ measurements

The full potential of an oceanic OCS and CS₂ database can be exploited, if measured concentrations are stored together with relevant metadata. As a minimal requirement, we recommend to report i) the exact date of each measurement, including time
10 of the day and ii) the exact location (including latitude, longitude and sample depth). This is especially important due to the photochemical production of both gases, as concentration in seawater varies strongly on diurnal and seasonal scales. To obtain a full diurnal cycle, we recommend to measure at least in a 4-hr-resolution, to minimize errors when interpolating and averaging over the period of one day. Of secondary importance are physical parameters such as temperature, radiation and wind speed. When modelling marine concentrations of OCS and CS₂, it is helpful to have access to the CDOM absorbance
15 data at a wavelength of 350 nm, because parameterizations for production rates are based on this value (von Hobe et al., 2003; Lennartz et al., 2017).

5 Data availability

The data is available from the PANGAEA database (<https://doi.pangaea.de/10.1594/PANGAEA.905430>, (Lennartz et al., 2019)).

20

6 Competing interests

The authors declare no competing interests.

Author contributions

25 S.T.L. and C.A.M. started the initiative for data compilation and long-term storage. All coauthors contributed data from one or more expeditions. S.T.L. wrote the manuscript with input from all coauthors.

References

Aranami, K.: Seasonal and regional comparison of oceanic and atmospheric dimethylsulfide in the northern north pacific: Dilution effects on its concentration during winter, *Journal of Geophysical Research*, 109, 10.1029/2003jd004288, 2004.



- Belviso, S., Mihalopoulos, N., and Nguyen, B. C.: The supersaturation of carbonyl sulfide (ocs) in rain waters, *Atmospheric Environment* (1987), 21, 1363-1367, 1987.
- 5 Berkelhammer, M., Steen-Larsen, H. C., Cosgrove, A., Peters, A., Johnson, R., Hayden, M., and Montzka, S. A.: Radiation and atmospheric circulation controls on carbonyl sulfide concentrations in the marine boundary layer, *Journal of Geophysical Research: Atmospheres*, 121, 2016.
- 10 Brühl, C., Lelieveld, J., Crutzen, P. J., and Tost, H.: The role of carbonyl sulphide as a source of stratospheric sulphate aerosol and its impact on climate, *Atmos. Chem. Phys.*, 12, 1239-1253, 10.5194/acp-12-1239-2012, 2012.
- Chin, M., and Davis, D. D.: Global sources and sinks of ocs and cs₂ and their distributions, *Glob. Biogeochem. Cycle*, 7, 321-337, 10.1029/93gb00568, 1993.
- 15 Crutzen, P. J.: The possible importance of cso for the sulfate layer of the stratosphere, *Geophysical Research Letters*, 3, 73-76, 10.1029/GL003i002p00073, 1976.
- Cutter, G. A., and Radford-Knoery, J.: Carbonyl sulfide in two estuaries and shelf waters of the western north atlantic ocean, *Marine Chemistry*, 43, 225-233, [http://dx.doi.org/10.1016/0304-4203\(93\)90228-G](http://dx.doi.org/10.1016/0304-4203(93)90228-G), 1993.
- 20 Cutter, G. A., Cutter, L. S., and Filippino, K. C.: Sources and cycling of carbonyl sulfide in the sargasso sea, *Limnology and Oceanography*, 49, 555-565, 2004.
- 25 de Gouw, J. A., Warneke, C., Montzka, S. A., Holloway, J. S., Parrish, D. D., Fehsenfeld, F. C., Atlas, E. L., Weber, R. J., and Flocke, F. M.: Carbonyl sulfide as an inverse tracer for biogenic organic carbon in gas and aerosol phases, *Geophysical Research Letters*, 36, n/a-n/a, 10.1029/2008gl036910, 2009.
- 30 Elliott, S., Lu, E., and Rowland, F. S.: Rates and mechanisms for the hydrolysis of carbonyl sulfide in natural waters, *Environmental Science & Technology*, 23, 458-461, 10.1021/es00181a011, 1989.
- Ferek, R. J., and Andreae, M. O.: The supersaturation of carbonyl sulfide in surface waters of the pacific oceans off peru, *Geophysical Research Letters*, 10, 393-395, 1983.
- 35 Ferek, R. J., and Andreae, M. O.: Photochemical production of carbonyl sulphide in marine surface waters, *Letters to Nature*, 1984, 148-150, 1984.
- Fischer, T., Kock, A., Arévalo-Martínez, D. L., Dengler, M., Brandt, P., and Bange, H. W.: Gas exchange estimates in the peruvian upwelling regime biased by multi-day near-surface stratification, *Biogeosciences Discuss.*, 2018, 1-34, 10.5194/bg-2018-395, 2018.
- 40 Inomata, Y., Matsunaga, K., Murai, Y., Osada, K., and Iwasaka, Y.: Simultaneous measurement of volatile sulfur compounds using ascorbic acid for oxidant removal and gas chromatography–flame photometric detection, *Journal of Chromatography A*, 864, 111-119, 1999.
- 45 Inomata, Y., Hayashi, M., Osada, K., and Iwasaka, Y.: Spatial distributions of volatile sulfur compounds in surface seawater and overlying atmosphere in the northwestern pacific ocean, eastern indian ocean, and southern ocean, *Glob. Biogeochem. Cycle*, 20, n/a-n/a, 10.1029/2005gb002518, 2006.
- Johnson, J. E.: Role of the oceans in the atmospheric cycle of carbonyl sulfide, PhD thesis, 1985.
- 50 Johnson, J. E., and Harrison, H.: Carbonyl sulfide concentrations in the surface waters and above the pacific ocean, *Journal of Geophysical Research*, 91, 7883-7888, 1986.
- Kamyshtny, A., Goifman, A., Rizkov, D., and Lev, O.: Formation of carbonyl sulfide by the reaction of carbon monoxide and inorganic polysulfides, *Environmental Science & Technology*, 37, 1865-1872, 10.1021/es0201911, 2003.
- 55



- Kettle, A. J., Rhee, T. S., von Hobe, M., Poulton, A., Aiken, J., and Andreae, M. O.: Assessing the flux of different volatile sulfur gases from the ocean to the atmosphere, *Journal of Geophysical Research: Atmospheres*, 106, 12193-12209, 10.1029/2000jd900630, 2001.
- Kim, K. H., and Andreae, M.: Carbon disulfide in seawater and the marine atmosphere over the north atlantic, *Journal of Geophysical Research: Atmospheres*, 92, 14733-14738, 1987.
- Kim, K. H., and Andreae, M. O.: Carbon disulfide in the estuarine, coastal and oceanic environments, *Marine Chemistry*, 40, 179-197, 1992.
- 10 Kremser, S., Jones, N. B., Palm, M., Lejeune, B., Wang, Y., Smale, D., and Deutscher, N. M.: Positive trends in southern hemisphere carbonyl sulfide, *Geophysical Research Letters*, 42, 9473-9480, 10.1002/2015gl065879, 2015.
- Kremser, S., Thomason, L. W., von Hobe, M., Hermann, M., Deshler, T., Timmreck, C., Toohey, M., Stenke, A., Schwarz, J. P., Weigel, R., Fueglistaler, S., Prata, F. J., Vernier, J.-P., Schlager, H., Barnes, J. E., Antuña-Marrero, J.-C., Fairlie, D., Palm, M., Mahieu, E., 15 Notholt, J., Rex, M., Bingen, C., Vanhellefont, F., Bourassa, A., Plane, J. M. C., Klocke, D., Carn, S. A., Clarisse, L., Trickl, T., Neely, R., James, A. D., Rieger, L., Wilson, J. C., and Meland, B.: Stratospheric aerosol—observations, processes, and impact on climate, *Reviews of Geophysics*, 54, n/a-n/a, 10.1002/2015rg000511, 2016.
- Launois, T., Belviso, S., Bopp, L., Fichot, C. G., and Peylin, P.: A new model for the global biogeochemical cycle of carbonyl sulfide - part 1: Assessment of direct marine emissions with an oceanic general circulation and biogeochemistry model, *Atmos. Chem. Phys.*, 15, 2295-2312, 10.5194/acp-15-2295-2015, 2015.
- 20 Lennartz, S. T., Marandino, C. A., von Hobe, M., Cortes, P., Quack, B., Simo, R., Booge, D., Pozzer, A., Steinhoff, T., Arevalo-Martinez, D. L., Kloss, C., Bracher, A., Röttgers, R., Atlas, E., and Krüger, K.: Direct oceanic emissions unlikely to account for the missing source of atmospheric carbonyl sulfide, *Atmos. Chem. Phys.*, 17, 385-402, 10.5194/acp-17-385-2017, 2017.
- 25 Lennartz, S. T., Marandino, C. A., von Hobe, M., Andreae, M. O., Aranami, K., Atlas, E. L., Berkelhammer, M., Bingemer, H., Booge, D., Cutter, G. A., Cortes, P., Kremser, S., Law, C. S., Marriner, A., Simo, R., Quack, B., Uher, G., Xie, H., and Xu, X.: A database for carbonyl sulfide (ocs) and carbon disulfide (cs₂) in seawater and marine boundary layer., in, PANGAEA, 30 <https://doi.pangaea.de/10.1594/PANGAEA.905430>, 2019.
- Mihalopoulos, N., Putaud, J. P., Nguyen, B. C., and Belviso, S.: Annual variation of atmospheric carbonyl sulfide in the marine atmosphere in the southern indian ocean, *Journal of Atmospheric Chemistry*, 13, 73-82, 10.1007/bf00048101, 1991.
- 35 Mihalopoulos, N., Nguyen, B. C., Putaud, J. P., and Belviso, S.: The oceanic source of carbonyl sulfide (cos), *Atmospheric Environment*, 26A, 1383-1394, 1992.
- Montzka, S.: Noaa whole air sampler mass spectrometer analysis.
Version 1.0. Ucar/ncar - earth observing laboratory, Global Monitoring Division (GMD), Earth System Research 40 Laboratory, 2013.
- Montzka, S. A.: A 350-year atmospheric history for carbonyl sulfide inferred from antarctic firn air and air trapped in ice, *Journal of Geophysical Research*, 109, 10.1029/2004jd004686, 2004.
- 45 Montzka, S. A., Calvert, P., Hall, B. D., Elkins, J. W., Conway, T. J., Tans, P. P., and Sweeney, C.: On the global distribution, seasonality, and budget of atmospheric carbonyl sulfide (cos) and some similarities to co₂, *Journal of Geophysical Research*, 112, 10.1029/2006jd007665, 2007.
- Moore, R., and Webb, M.: The relationship between methyl bromide and chlorophyll α in high latitude ocean waters, *Geophysical Research Letters*, 23, 2951-2954, 1996.
- 50 Radford-Knoery, J., and Cutter, G. A.: Biogeochemistry of dissolved hydrogen sulfide species and carbonyl sulfide in the western north atlantic ocean, *Geochimica et Cosmochimica Acta*, 58, 5421-5431, [http://dx.doi.org/10.1016/0016-7037\(94\)90239-9](http://dx.doi.org/10.1016/0016-7037(94)90239-9), 1994.
- 55 Staubes, R., Georgii, H. W., and Ockelmann, G.: Flux of cos, dms and cs₂ from various soils in germany, *Tellus B*, 41, 305-313, 1989.



- Staubes, R., Georgii, H.-W., and Bürgermeister, S.: Biogenic sulfur compounds in seawater and the marine atmosphere, in: Physico-chemical behaviour of atmospheric pollutants, Springer, 686-692, 1990.
- 5 Staubes, R., and Georgii, H.-W.: Biogenic sulfur compounds in seawater and the atmosphere of the antarctic region, *Tellus*, 45B, 127-137, 1993.
- Uher, G.: Photochemische produktion von carbonylsulfid im oberflächenwasser der ozeane: Prozeßstudien und ein empirisches modell, Universität Mainz, 147 pp., 1994.
- 10 Uher, G., and Andreae, M. O.: Photochemical production of carbonyl sulfide in north sea water: A process study, *Limnology and Oceanography*, 42, 432-442, 1997.
- Ulshöfer, V. S., Uher, G., and Andreae, M. O.: Evidence for a winter sink of atmospheric carbonyl sulfide in the northeast atlantic ocean, *Geophysical Research Letters*, 22, 2601-2604, 1995.
- 15 Ulshöfer, V. S., Flöck, O. R., Uher, G., and Andreae, M. O.: Photochemical production and air-sea exchange of carbonyl sulfide in the eastern mediterranean sea, *Marine Chemistry*, 53, 25-39, 1996.
- Ulshöfer, V. S., and Andreae, M. O.: Carbonyl sulfide (cos) in the surface ocean and the atmospheric cos budget, *Atmospheric Geochemistry*, 3, 283-303, 1998.
- 20 Von Hobe, M., Kettle, A. J., and Andreae, M. O.: Carbonyl sulphide in and over seawater: Summer data from the northeast atlantic ocean, *Atmospheric Environment*, 33, 3503-3514, 1999.
- 25 von Hobe, M., Kenntner, T., Helleis, F. H., Sandoval-Soto, L., and Andreae, M. O.: Cryogenic trapping of carbonyl sulfide without using expendable cryogenes, *Analytical chemistry*, 72, 5513-5515, 2000.
- Von Hobe, M., Cutter, G. A., Kettle, A. J., and Andreae, M. O.: Dark production: A significant source of oceanic cos, *Journal of Geophysical Research*, 106, 31217, 10.1029/2000jc000567, 2001.
- 30 von Hobe, M., Najjar, R. G., Kettle, A. J., and Andreae, M. O.: Photochemical and physical modeling of carbonyl sulfide in the ocean, *Journal of Geophysical Research*, 108, 10.1029/2000jc000712, 2003.
- Von Hobe, M., Kuhn, U., Van Diest, H., Sandoval-Soto, L., Kenntner, T., Helleis, F., Yonemura, S., Andreae, M. O., and Kesselmeier, J.: Automated in situ analysis of volatile sulfur gases using a sulfur gas analyser (sugar) based on cryogenic trapping and gas-chromatographic separation, *International Journal of Environmental Analytical Chemistry*, 88, 303-315, 10.1080/03067310701642081, 2008.
- 35 Watts, S. F.: The mass budgets of carbonyl sulfide, dimethyl sulfide, carbon disulfide and hydrogen sulfide, *Atmospheric Environment*, 34, 761-779, 10.1016/s1352-2310(99)00342-8, 2000.
- 40 Weiss, P. S., Andrews, S. S., Johnson, J. E., and Zafiriou, O. C.: Photoproduction of carbonyl sulfide in south pacific ocean waters as a function of irradiation wavelength, *Geophysical Research Letters*, 22, 215-218, 1995a.
- 45 Weiss, P. S., Johnson, J. E., Gammon, R. H., and Bates, T. S.: Reevaluation of the open ocean source of carbonyl sulfide to the atmosphere, *Journal of Geophysical Research*, 100, 23083-23092, 1995b.
- Whelan, M. E., Lennartz, S. T., Gimeno, T. E., Wehr, R., Wohlfahrt, G., Wang, Y., Kooijmans, L. M. J., Hilton, T. W., Belviso, S., Peylin, P., Commane, R., Sun, W., Chen, H., Kuai, L., Mammarella, I., Maseyk, K., Berkelhammer, M., Li, K. F., Yakir, D., Zumkehr, A., Katayama, Y., Ogée, J., Spielmann, F. M., Kitz, F., Rastogi, B., Kesselmeier, J., Marshall, J., Erkkilä, K. M., Wingate, L., Meredith, L. K., He, W., Bunk, R., Launois, T., Vesala, T., Schmidt, J. A., Fichot, C. G., Seibt, U., Saleska, S., Saltzman, E. S., Montzka, S. A., Berry, J. A., and Campbell, J. E.: Reviews and syntheses: Carbonyl sulfide as a multi-scale tracer for carbon and water cycles, *Biogeosciences*, 15, 3625-3657, 10.5194/bg-15-3625-2018, 2018.
- 55 Wofsy, S. C., Afshar, S., Allen, H. M., Apel, E., Asher, E. C., Barletta, B., Bent, J., Bian, H., Biggs, B. C., Blake, D. R., Blake, N., Bourgeois, I., Brock, C. A., Brune, W. H., Budney, J. W., Bui, T. P., Butler, A., Campuzano-Jost, P., Chang, C. S., Chin, M., Commane,



- R., Correa, G., Crounse, J. D., Cullis, P. D., Daube, B. C., Day, D. A., Dean-Day, J. M., Dibb, J. E., Digangi, J. P., Diskin, G. S., Dollner, M., Elkins, J. W., Erdesz, F., Fiore, A. M., Flynn, C. M., Froyd, K., Gesler, D. W., Hall, S. R., Hanisco, T. F., Hannun, R. A., Hills, A. J., Hints, E. J., Hoffman, A., Hornbrook, R. S., Huey, L. G., Hughes, S., Jimenez, J. L., Johnson, B. J., Katich, J. M., Keeling, R., Kim, M. J., Kupc, A., Lait, L. R., Lamarque, J. F., Liu, J., McKain, K., McLaughlin, R. J., Meinardi, S., Miller, D. O., Montzka, S. A., Moore, F. L., Morgan, E. J., Murphy, D. M., Murray, L. T., Nault, B. A., Neuman, J. A., Newman, P. A., Nicely, J. M., Pan, X., Paplawsky, W., Peischl, J., Prather, M. J., Price, D. J., Ray, E., Reeves, J. M., Richardson, M., Rollins, A. W., Rosenlof, K. H., Ryerson, T. B., Scheuer, E., Schill, G. P., Schroder, J. C., Schwarz, J. P., St.Clair, J. M., Steenrod, S. D., Stephens, B. B., Strode, S. A., Sweeney, C., Tanner, D., Teng, A. P., Thames, A. B., Thompson, C. R., Ullmann, K., Veres, P. R., Vizenor, N., Wagner, N. L., Watt, A., Weber, R., Weinzierl, B., Wennberg, P., Williamson, C. J., Wilson, J. C., Wolfe, G. M., Woods, C. T., and Zeng, L. H.: Atom: Merged atmospheric chemistry, trace gases, and aerosols, in, ORNL Distributed Active Archive Center, 2018.
- Xie, H., and Moore, R. M.: Carbon disulfide in the north atlantic and pacific oceans, *Journal of Geophysical Research*, 104, 5393, 10.1029/1998jc900074, 1999.
- Xu, X., Bingemer, H. G., Georgii, H. W., Schmidt, U., and Bartell, U.: Measurements of carbonyl sulfide (cos) in surface seawater and marine air, and estimates of the air-sea flux from observations during two atlantic cruises, *Journal of Geophysical Research*, 106, 3491, 10.1029/2000jd900571, 2001.
- Zhang, L., Walsh, R. S., and Cutter, G. A.: Estuarine cycling of carbonyl sulfide: Production and sea-air flux, *Marine Chemistry*, 61, 127-142, 1998.



Tables

Table 1. Flags used to describe sampling frequency of each individual dataset.

0	not reported
1	minutely
2	15-minutely
3	hourly
4	1-4 hourly
5	>4 hourly to daily
6	monthly
7	seasonally
8	annually
9	irregular

5

10 **Table 2. Description of all cruises or campaigns contributing measurements to this database. Cruises are given a unique ID for identification. Reference refers to the publication where the data was reported first. Methods are reported using the same specifications and level of detail as given in the original publication. Specifications for analytical methods are listed together with the method referenced in the main reference. d/o stands for digitalized/original data. S=sampling, A=analysis, D=details, R=reference of instrumentation. Letters in the last column indicate direct comparability of the datasets, as studies were performed with either identical analytical systems (i.e. same method reference), or performed by the same laboratory through further development of analytical systems (i.e. different method reference, but intercomparison within laboratory).**

ID	campaign/ship/ date/region	reference	method	data source	grouping
1	RV Robert Conrad June 1982 ETSP	Ferek and Andreae (1983)	OCS S: gas bubbler in glass column A: GC-FPD D: precision <10%, st. dev. of triplicates 6% R: Ferek and Andreae (1983),	d	A
2	RV Cape Hatteras April 1983 Chesapeake Bay	Ferek and Andreae (1984)	OCS S: gas bubbler in glass column A: GC-FPD R: Ferek and Andreae (1983)	d	A
3	RV Discoverer March-June 1982	Johnson and Harrison (1986)	OCS S: glass syringes/ bucket	d	B



	Pacific		A: GC-FPD D: standard from permeation tubes, l.o.d. 0.04 ngS, reproducibility within 2.5%, not georeferenced R: Johnson (1985)		
4	RV Discoverer March-June 1983 Pacific	Johnson and Harrison (1986)	OCS S: glass syringes/ bucket A: GC-FPD D: standard from permeation tubes, l.o.d. 0.04 ngS, reproducibility within 2.5%, not georeferenced R: Johnson (1985)	d	B
5	RV Columbus Iselin April-May 1986 Atlantic	Kim and Andreae (1992)	CS ₂ S: purged with N ₂ A: GC-FPD D: liquid CS ₂ standard cross checked with gas standard Metronic Associates Inc.(Santa Clara, CA), precision 9%, l.o.d. 2 pmol S L ⁻¹ R: Kim and Andreae (1987)	d	C
6	Coastal 1987-1888 Amsterdam Island, Indian Ocean	Mihalopoulos et al. (1991)	OCS S: pressurized electropolished stainless steel canisters A:GC-FPD/FPD D: standard: gas standard Matheson Union Carbide, l.o.d. 0.4 ng OCS = 53 ppt, reproducibility <5% (8 repeats), accuracy 10%, not fully georeferenced R: Belviso et al. (1987)	d	D
7	RV Polarstern 1988 Atlantic, Southern Ocean	Staubes et al. (1990)	OCS, CS ₂ S: purged with N ₂ A: GC-FPD D: l.o.d. 3-4pptv and 0.5-1 ngS L ⁻¹ , not georeferenced R: Staubes et al. (1990)	d	E
8	RV Cape Hatteras April 1989 North Atlantic	Radford-Knoery and Cutter (1994)	OCS S: sampling with Go-Flo bottles, acidified, stripped with He A: GC-FPD D: l.o.d. 1.3 pmol L ⁻¹ , precision 5% R: Cutter and Radford-Knoery (1993)	d	F
9	RV Cape Hatteras November 1989 Estuary, North Atlantic	Radford-Knoery and Cutter (1994)	OCS S: sampling with Go-Flo bottles, acidified, stripped with He A: GC-FPD D: l.o.d. 1.3 pmol L ⁻¹ , precision 5% R: Cutter and Radford-Knoery (1993)	d	F
10	RV Cape Henlopen June 1990 North Atlantic	Cutter and Radford-Knoery (1993)	OCS S: Go-Flo bottles, gas-tight syringes, stripped with He A: GC-FPD D: l.o.d. 1,3 pmol/L, precision 5%, R: Cutter and Radford-Knoery (1993)	d	F



11	RV Polarstern November 1990 Southern Ocean	Staubes and Georgii (1993)	OCS S: air: directly to sample loop, water: into gas stripping column with N ₂ A: GC-FPD D: l.o.d. 3.5ppt, 6.4% precision R: Staubes et al. (1989)	d	E
12	OCEAT II+III, diverse 1987-1991 Mediterranean Sea, Red Sea, Indian Ocean	Mihalopoulos et al. (1992)	OCS S: pressurized electropolished stainless steel canisters A: GC-FPD D: lod: 0.4 ng S, precision 10%, not georeferenced R: Mihalopoulos et al. (1992)	d	D
13	Chesapeake Bay time series 1991-1994 Chesapeake Bay	Zhang et al. (1998)	OCS S: depth profiles with pump, Go-Flo, cubitainer, stripped with He A: GC-FPD C: standard: permeation tubes, precision <5%, l.o.d. 10 pmol OCS L ⁻¹ R: Cutter and Radford-Knoery (1993)	d	F
14	RV Meteor – M21 April 1992 North Atlantic, North Sea	Ulshöfer et al. (1995)	OCS S: equilibrator A: GC-FPD D: standard: permeation tubes, l.o.d. OCS: 100ppt, reproducibility 15% R: Uher (1994)	o	G
15	FP Nordsee September 1992 North Sea	Uher and Andreae (1997)	OCS S: equilibrator A: GC-FPD D: standard: permeation tubes, l.o.d. 100 pg OCS=105ppt, precision <15% R: Ulshöfer et al. (1995), (Uher, 1994)	o	G
16	RV Aegaio, EGAMES July 1993 Mediterranean Sea	Ulshöfer et al. (1996)	OCS S: equilibrator A: GC-FPD D: standard: permeation tubes, detection limit OCS: 4 pmol L ⁻¹ , precision 15% R: Uher (1994)	o	G
17	RV Surveyor November 1993 Pacific	Weiss et al. (1995b)	OCS S: glass syringe, purge-and-trap A: GC-ECD D: standard: permeation tubes, l.o.d. 115 ppt cruise1, 23ppt cruise 2, uncertainty 6-10%, reproducibility of blanks 7% R: Weiss et al. (1995a)	d	B
18	RV Meteor – M27 January 1994 North Atlantic, North Sea	Ulshöfer et al. (1995)	OCS S: equilibrator A: GC-FPD D: standard: permeation tubes, l.o.d. OCS: 100 ppt, reproducibility 10% R: Uher (1994)	o	G
19	RV Valdivia April 1994	Ulshöfer and Andreae (1998)	OCS S: Weiss-type equilibrator	o	G



	North Sea		A: GC-FPD D: standard from permeation tubes, precision <10% R: Ulshöfer et al. (1995), (Uher, 1994)		
20	RV Columbus Iselin August 1994 North Atlantic (Florida)	Ulshöfer and Andreae (1998)	OCS S: Weiss-type equilibrator A: GC-FPD D: standard from permeation tubes, precision <10% R: Ulshöfer et al. (1995), (Uher, 1994)	o	G
21	RV Meteor – M30 September 1994 North Atlantic, North Sea	Ulshöfer et al. (1995)	OCS S: equilibrator A: GC-FPD D: standard: permeation tubes, l.o.d. 100ppt, reproducibility <10% R: Ulshöfer et al. (1995)	o	G
22	RV Cape Hatteras March 1995 North Atlantic, Bermuda	Ulshöfer and Andreae (1998)	OCS S: Weiss-type equilibrator A: GC-FPD D: standard from permeation tubes, precision <10% R: Ulshöfer et al. (1995)	o	G
23	RV Hudson July 1995 Atlantic, Pacific	Xie and Moore (1999)	CS ₂ S: bucket/submersible pump A: GC-MS D: gravimetrically prepared liquid standard, l.o.d. CS ₂ 1.5 pmol L ⁻¹ S, rel.st.dev. 1.4% at 10 pmol L ⁻¹ level R: Moore and Webb (1996)	o	H
24	RV Discoverer October 1995 Atlantic, Pacific	Xie and Moore (1999)	CS ₂ S: stainless-steel Knudsen bottles A: GC-MS D: gravimetrically prepared liquid standard, l.o.d. CS ₂ 1.5 pmol L ⁻¹ S, rel.st.dev. 1.4% at 10 pmol L ⁻¹ level R: Moore and Webb (1996)	o	H
25	RV Shirase November 1996 Indian Ocean, Southern Ocean	Inomata et al. (2006)	OCS S: PTFE-tubing, Flek-sampler A: GC-FPD D: standard gas (Nippon Sanso Co. Ltd.), l.o.d. 0.06nmol L ⁻¹ /12ppt, uncertainty 6% R: Inomata et al. (1999)	d	I
26	RV Prof. Vodyanitsky, ACE-2 June 1997 North Atlantic	Von Hobe et al. (1999)	OCS S: Weiss-equilibrator A: GC-FPD, D: standard: permeation tubes, l.o.d. 30ppt/0.4 pmol/L, precision <10% R: Von Hobe et al. (1999)	o	G
27	KH97-2 Jul 1997 North Pacific	Aranami (2004)	OCS S: Tedlar-bags A: GC-FPD D: gas cylinder standard: Takachiho Kogyo Co. Ltd., precision 5%	o	J



			R: Aranami (2004)		
28	RV Polarstern ANTXV-1 November 1997 Atlantic transect	Xu et al. (2001)	OCS, CS ₂ S: equilibrator A: GC-FPD D: standard: permeation tubes, precision 3%, uncertainty 10% R: Xu et al. (2001)	o	K
29	RV Polarstern ANTXV-5 May 1998 Atlantic transect	Xu et al. (2001)	OCS, CS ₂ S: Teflon-equilibrator A: GC-FPD D: standard: permeation tubes, precision 3%, uncertainty 10% R: Xu et al. (2001)	o	K
30	RV Mirai MR98-K01 Nov. 1998 North Pacific	Aranami (2004)	OCS, CS ₂ S: air: Tedlar-bags, seawater: plastic bucket, glass syringe A: GC-FPD D: gas cylinder standard: Takachiho Kogyo Co. Ltd., precision 5% R: Aranami (2004)	o	J
31	RV Endeavor 327 April 1999 North Atlantic, BATS	Von Hobe et al. (2001)	OCS S:equilibrator A: GC-FPD D: precision <2%, standard from permeation tubes, lod: 30 ppt/0.4pmol OCS R: von Hobe et al. (2000)	o	G
32	BATS August 1999 North Atlantic, BATS	Cutter et al. (2004)	OCS S: Go-Flo bottles, submersible pumping system A: GC-FPD D: l.o.d. 1 pmol L ⁻¹ , precision <10% R: Radford-Knoery and Cutter (1994)	o	F
33	RV James Clark Ross, AMT-7 September 1999 Atlantic	Kettle et al. (2001)	OCS, CS ₂ S: equilibrator A: GC-FPD C: permeation tubes R: Ulshöfer et al. (1995)	d	G
34	RV Poseidon P269 February 2001 Atlantic Ocean	partially published in Von Hobe et al. (2008)	OCS, CS ₂ S:equilibrator A: GC-FPD D: precision 1.9 % for COS and 2.2 % for CS ₂ , standard from permeation tubes, lod: 20 ppt/0.3pmol OCS and 10 ppt CS ₂ R: Von Hobe et al. (2008)	o	G
35	RV Sonne, SHIVA November 2011 Pacific, Indian Ocean	unpublished	OCS S: gas canister A: GC-MS D: referenced to NOAA standard, precision 1%, calibration accuracy 10% R: de Gouw et al. (2009)	o	L
36	RV Sonne, SPACES-OASIS July 2014	Lennartz et al. (2017)	OCS S: equilibrator A: OA-ICOS	o	M



	Indian Ocean		D: standard permeation tubes, 15ppt precision, l.o.d. 4pmol/L \approx 200ppt standard within 2% of NOAA scale R: Lennartz et al. (2017)		
37	RV Atlantic Explorer September 2014 North Atlantic, BATS	Berkelhammer et al. (2016)	OCS S: air via tube to instrument A: OA-ICOS D: referenced against NOAA standard, std. dev. 12.7 ppt R: Berkelhammer et al. (2016)	o	N
38	RV Hesperides, TransPEGASO October 2014 Atlantic	Lennartz et al. (2017)	OCS, CS ₂ S: glass bottles A: GC-MS D: standards: gas (OCS), liquid (CS ₂), l.o.d. OCS: 1.8 pmol L ⁻¹ , CS ₂ : 1.4 pmol L ⁻¹ , precision \sim 5% R:Lennartz et al. (2017)	o	O
39	Tudor Hill Observatory Dec 2014-Mar 2015 Bermuda	Berkelhammer et al. (2016)	OCS S: air via tube to instrument A: OA-ICOS D: referenced against NOAA standard, std. dev. 12.7 ppt R: Berkelhammer et al. (2016)	o	N
40	RV Sonne, ASTRA- OMZ August 2015 ETSP	Lennartz et al. (2017)	OCS, CS ₂ S: OCS: equilibrator, CS ₂ : Niskin bottles A: OCS: OA-ICOS, CS ₂ : GC-MS D: OCS standard permeation tubes, 15ppt precision, l.o.d. 4pmol/L \approx 200ppt, standard within 2% of NOAA scale, CS ₂ : l.o.d. 1 pmol L ⁻¹ , precision 5-10% R: Lennartz et al. (2017)	o	M
41	RV Tangaroa Feb-Mar 2018 Southern Ocean	unpublished	OCS S: equilibrator A: OA-ICOS D: standard permeation tubes, 15ppt precision, l.o.d. 4pmol/L \approx 200ppt R:Lennartz et al. (2017)	o	M
42	RV Xue Long Nov 2018 – Mar 2019 Pacific, Southern Ocean	unpublished	OCS S: Spray-head equilibrator A: OA-ICOS D: standard permeation tubes, accuracy <18%, l.o.d. 4 pmol/L R: Lennartz et al. (2017)	o	M



Table 3: Quantitative sampling details for each individual dataset. #=number of samples, depth/height = water depth below sea surface, height above sea surface, t.r.=flag for temporal resolution, see Tab. 1. mbl=marine boundary layer, t.d.= top deck (height not specified).

ID	OCS seawater			OCS mbl			CS ₂ seawater			CS ₂ mbl		
	#	depth [m]	t.r.	#	height [m]	t.r.	#	depth [m]	t.r.	#	height [m]	t.r.
1	13	4	5									
2	18	1-18	5									
3	33	5	9	46	10	9						
4	115	5	9	101	10	9						
5							61	0-302	9			
6				8	25	0						
7	62		5	62	21	5	61	surface	5	58	21	5
8	18	16-850	5									
9	33	2-995	5									
10	5	3	5									
11	124	2	5	99	21	0						
12	109	diverse	0/6									
13	104	0-18	9									
14	118	3	3	118	28	3						
15	69	5	4	69	38	4						
16	123	3	4	123	6	4						
17				43		0						
18	120	7	4	120	38	4						
19	168	4	3	168	t.d.	3						
20	50	4	3	50	t.d.	3						
21	235	7	3	235	30	3						
22	323	4	3	323	t.d.	3						
23							17	surface	6			
24							17	surface	6			
25				12		5						
26	940	2-5	2	4175	10	2						
27				50	5	9				72		9
28	306	5	3	306	25	3	306	5	3	306	25	3
29	440	5	3	441	10-15	3	440	5	3	440	10-15	3
30	46	0-40	9	65	5	9	45		9	65	5	9
31	518		2	167		3						
32	132	1-300	4									
33	345	6	9	192		9	235	6	9			
34	287	1	2	95		3	287		2	95		3
35				193		4						
36	206	5	3	210	35	3						
37				1930		2						



38	42	5	5				42		5			
39				2213	23	3*						
40	285	5	3	256	35	3	144	5	3			
41	421		3	527		3						
42	1727	6	3	1908	t.d.	3						

*original paper in sampling frequency of seconds, averaged for this database



Figures

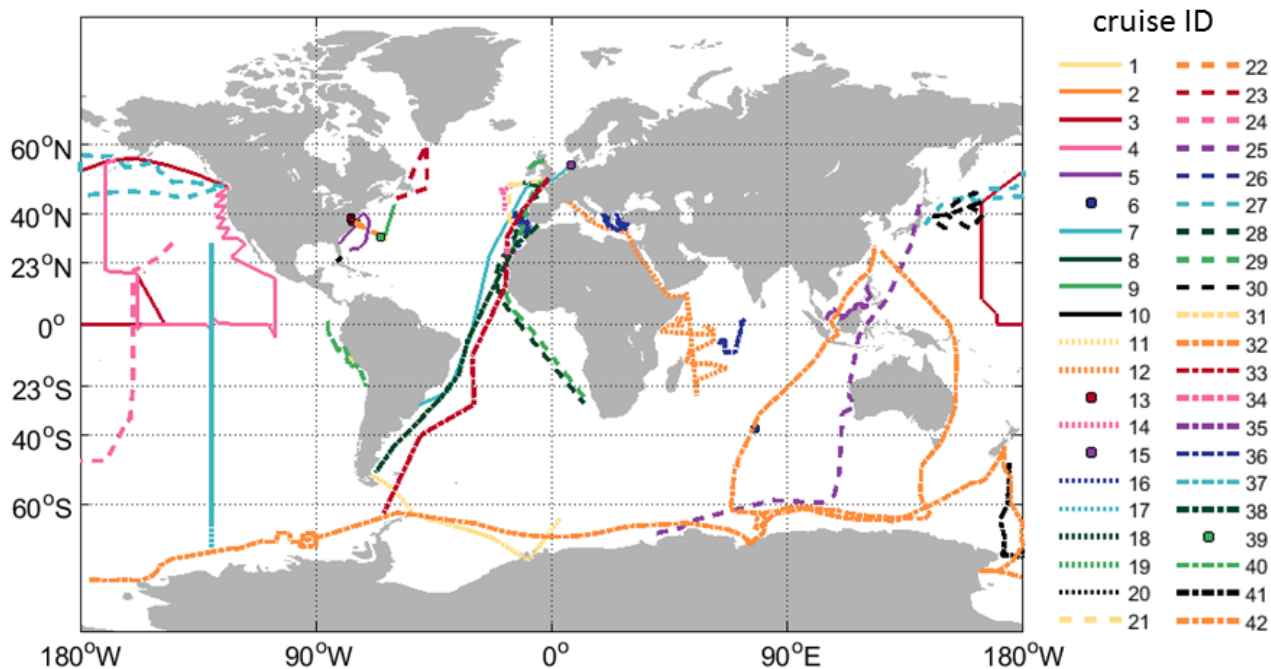


Figure 1: Tracks of all cruises with OCS and/or CS₂ measurements included in the database (points depict stationary measurements). Colour coding and line styles indicate the cruise ID (compare Tab. 1).

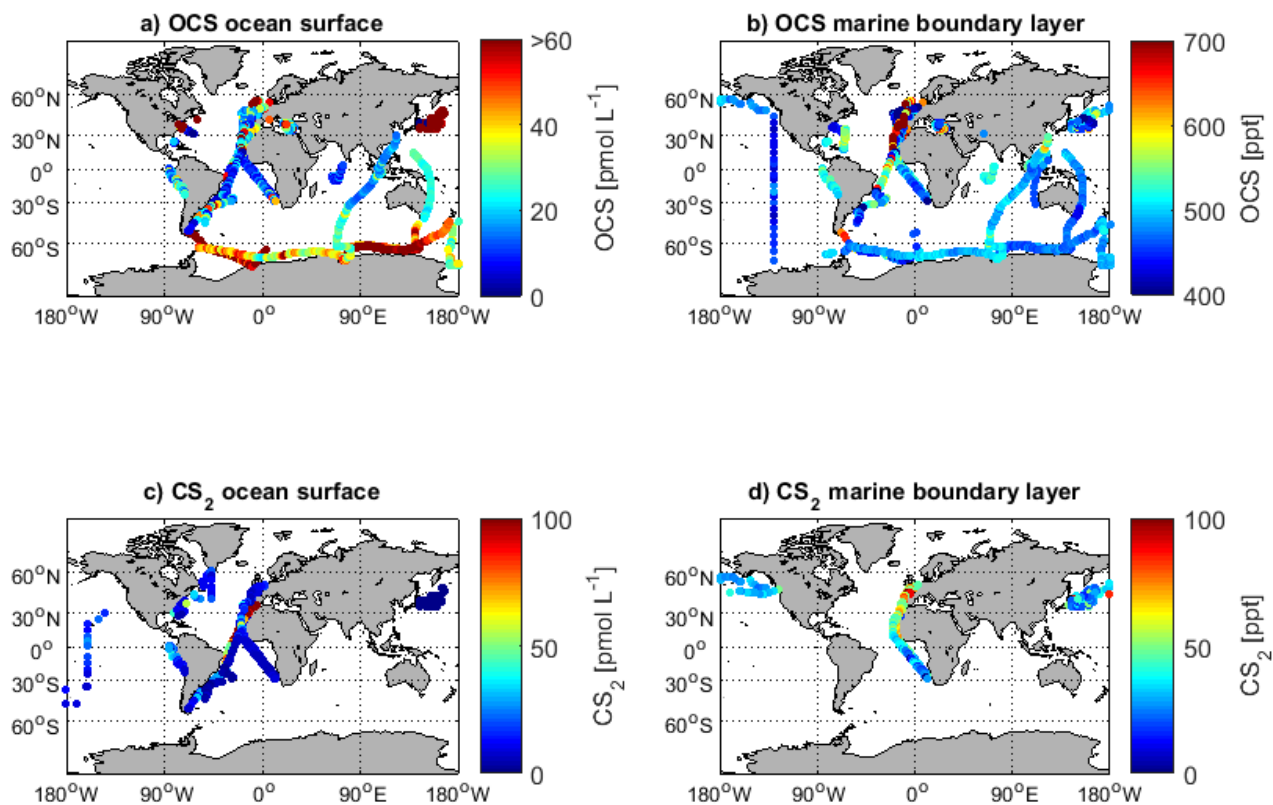
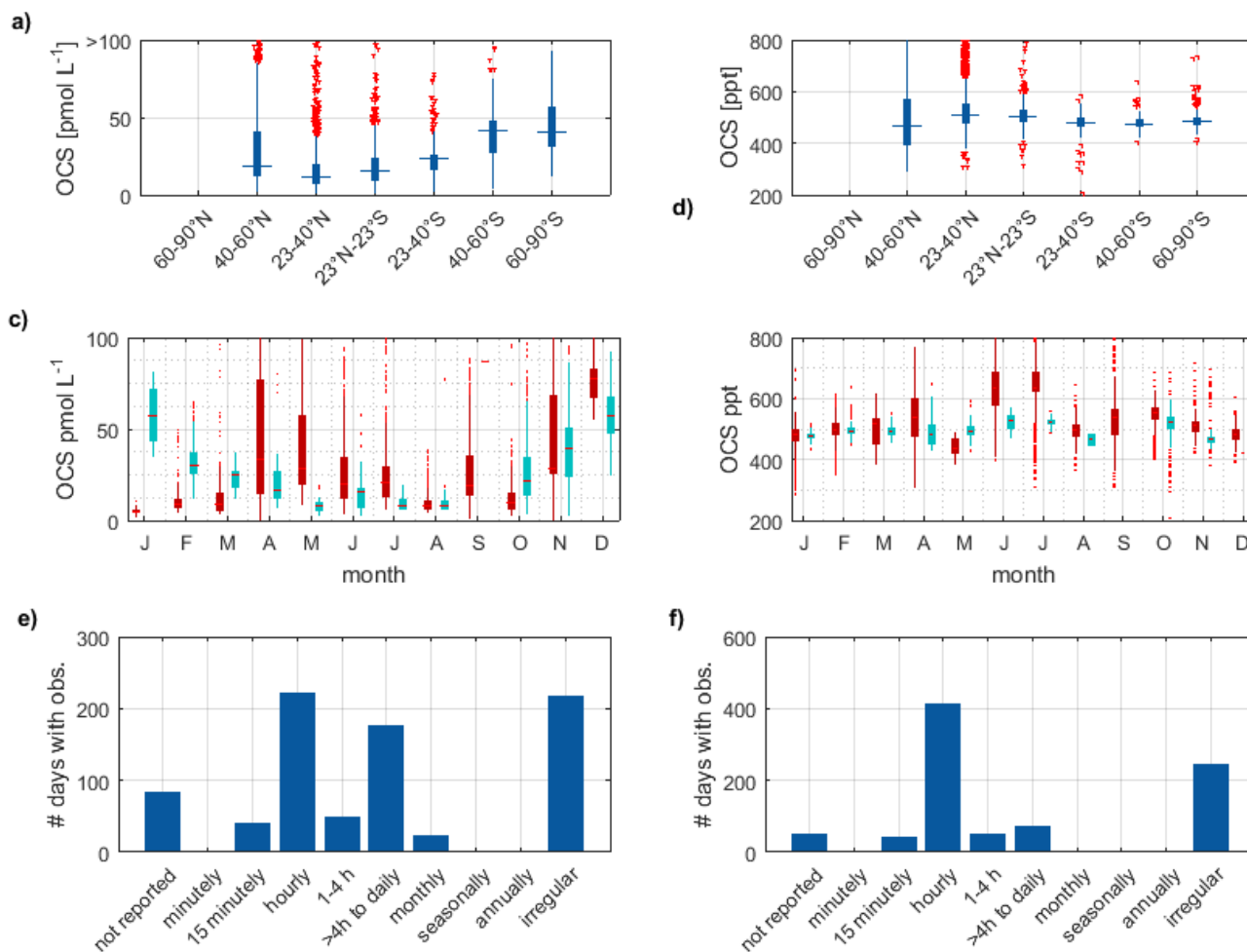


Figure 2: Georeferenced data for a) surface ocean OCS concentrations, b) marine boundary layer OCS mixing ratios, c) surface ocean CS₂ concentrations, d) marine boundary layer CS₂ mixing ratios. Only surface data (shallower than 10m) are shown.



5 **Figure 3: Overview of the OCS datasets: Boxplots of concentrations per latitudinal bin for a) water and b) marine boundary layer measurements. Blue boxes show range of 25 and 75 percentile, horizontal bar indicates the median, and red crosses show outliers. The temporal variation averaged over all years for c) water and d) marine boundary layer (note that in panel c and d, red indicates Northern hemisphere data, whereas light blue indicates Southern hemisphere data. Note that measurements >150 pmol L⁻¹ were excluded from these statistics (i.e. coastal samples). Numbers of days with observations for temporal resolution from minutely to annually for e) water and f) marine boundary layer measurements. (Note that for the boxplots in a) and b), only completely georeferenced data were included).**

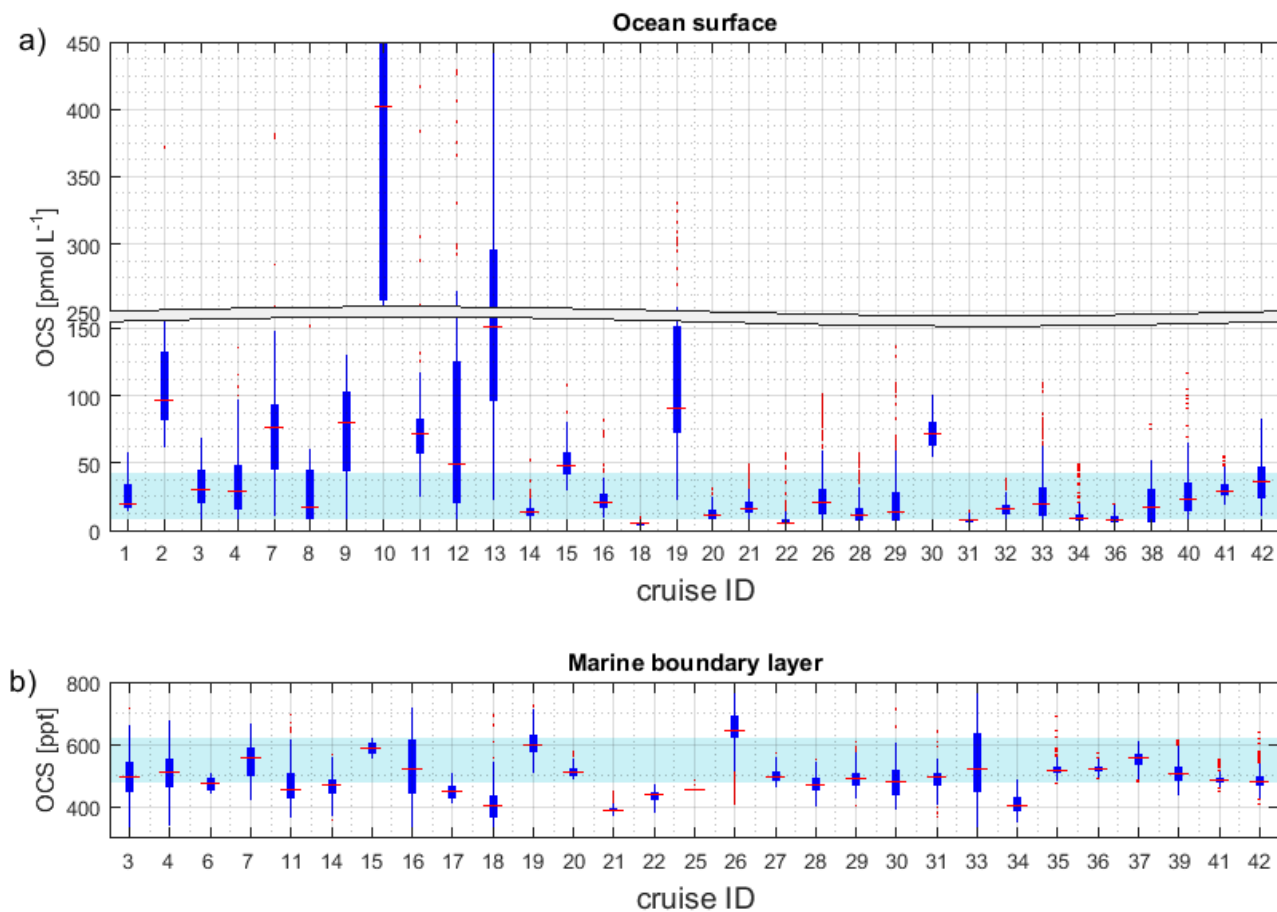


Figure 4: Boxplots of measured OCS concentrations in a) seawater and b) marine boundary layer. Marked in red is the median of each individual dataset, the edges of the box represent the 25th and 75th percentile, and outliers are indicated by red dots. The patch in the background indicates the 20th and 80th percentile of the whole dataset. Note the break in the y-axis in a).

5

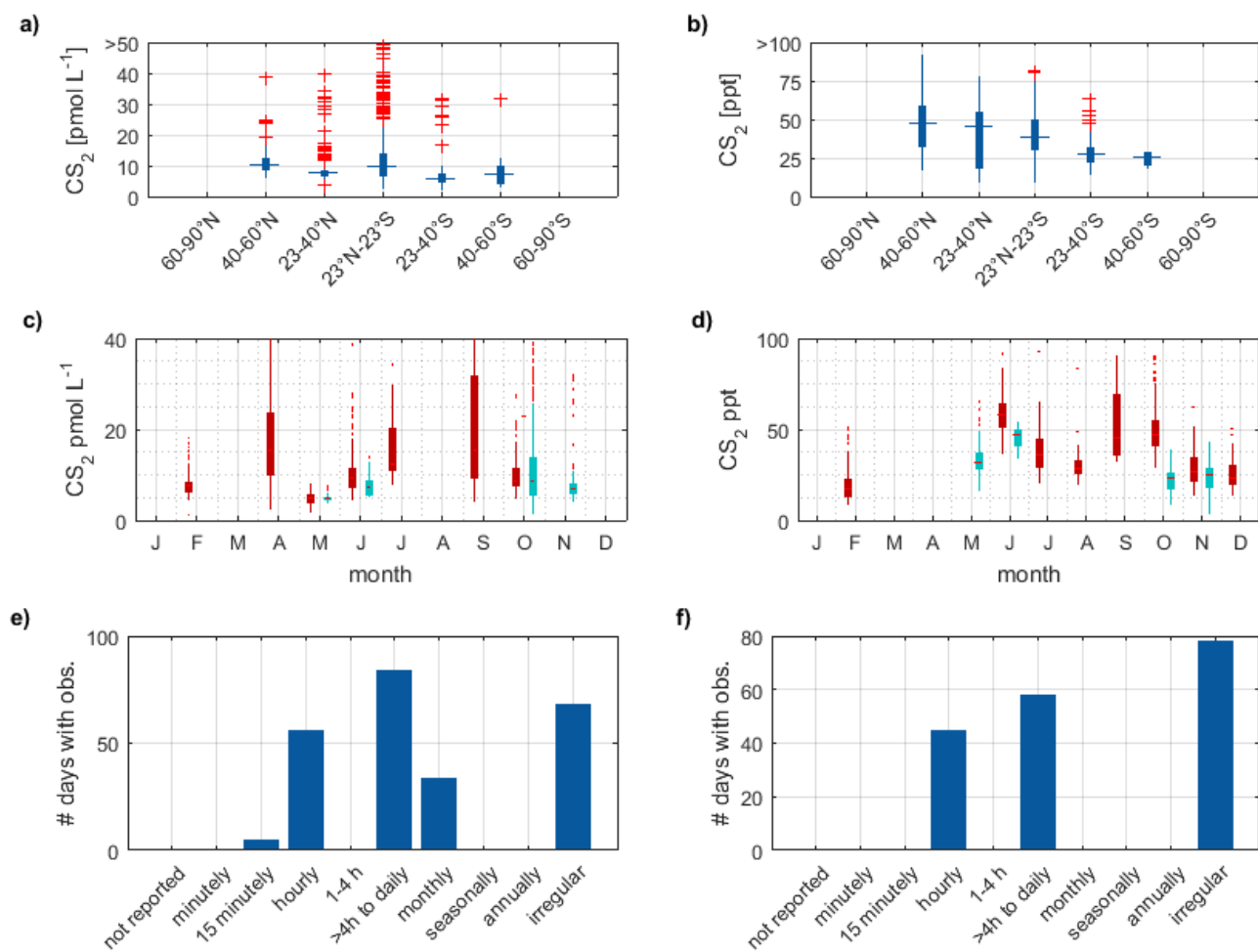


Figure 5: Same as Figure 2, but for CS₂.

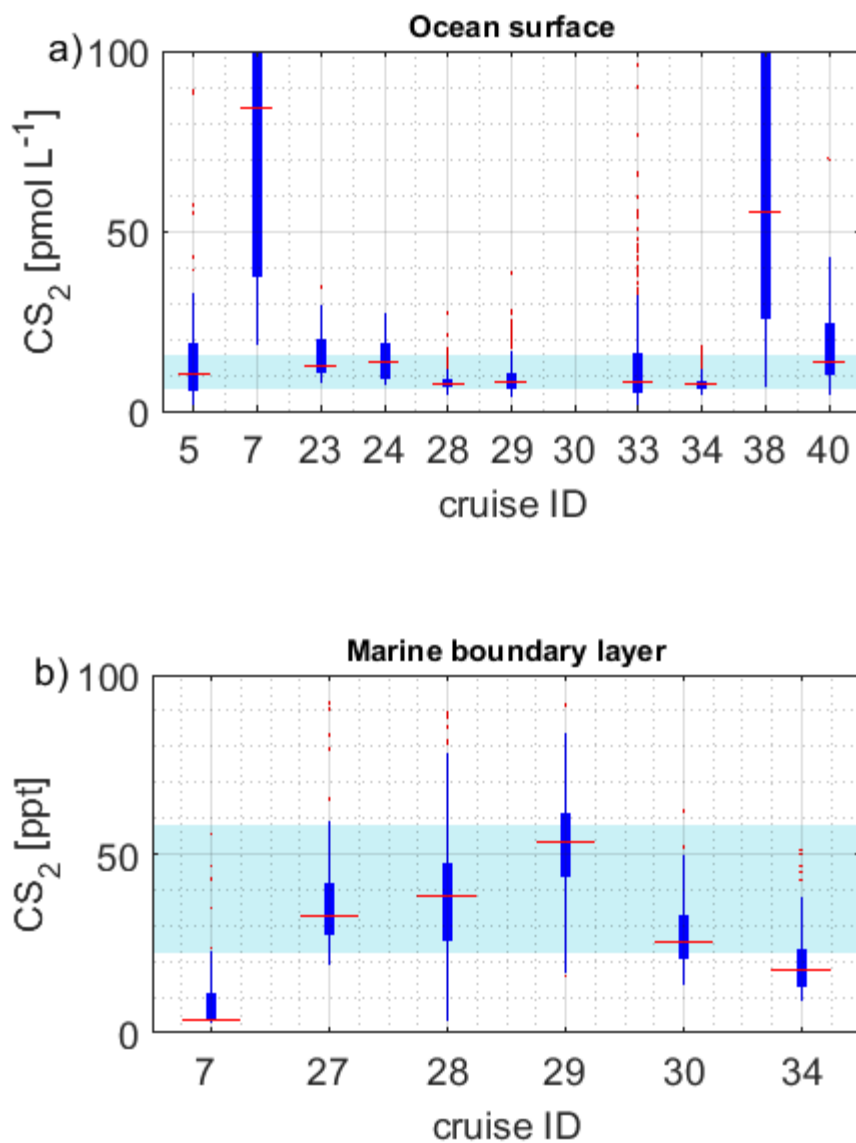


Figure 6 Same as Figure 4 but for CS₂.